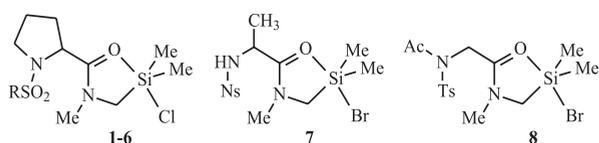


interactions in the structures **1-6** are the weak (CH ... O and CH ... Cl) bonds. In addition, in the structure **6** the stacking interactions between the fragments Ph-NO<sub>2</sub> are observed.

The derivatives of glycine and alanine (**7** and **8**) can be described as “freeze” cation-anion pairs. The Si...Br interatomic distances (2.7095(8) and 2.9704(7) Å, for **7** and **8**, respectively) are much larger than those in the compounds of tetracoordinated Si (for instance, in the tris(pentafluorophenylmethyl)silylbromide the corresponding bond length is 2.208 Å). In turn, the Si-O bond in **7** and **8** is considerably shorter than in **1-6**. In contrast to **1-6**, the coordination polyhedron corresponds to the distorted tetrahedron with weak additional Si-O coordination bond.

To estimate the energy of intermolecular interactions, we have carried out quantum chemical calculations of the isolated molecules **7-8** and their associates. Also, the solid-state MAS <sup>29</sup>Si spectra of **7** and **8** were simulated using plane-wave DFT calculations.

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1: R=Me, 2: R=Ph, 3: R=p-MeC<sub>6</sub>H<sub>4</sub>, 4: R=p-ClC<sub>6</sub>H<sub>4</sub>, 5: R=p-BrC<sub>6</sub>H<sub>4</sub>, 6: R=p-NO<sub>2</sub>C<sub>6</sub>H<sub>4</sub>

**Keywords:** N-organosulfonyl-2-aminoacids, hypervalent silicon, crystal packing

## MS61.P10

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**The influence of crystal packing on the structure of N-organosulfonyl-2-aminoacid complexes containing OSiC<sub>3</sub>F coordination unit**

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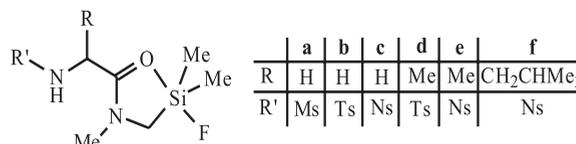
Silylfluorides with pentacoordinated Si atom obtained from N-organosulfonyl-2-aminoacids contain bioactive fragments and have high hydrolytic stability. In this study, we analyzed the influence of crystal packing on the geometry of coordination environment of the silicon atom in six pentacoordinated silylfluorides containing OSiC<sub>3</sub>F coordination polyhedron.

Taking into account similar inductive effect of exocyclic substituents in the complexes **a-f**, one would expect the coordination site of the silicon atom to have a similar structure. Indeed, in all cases OSiC<sub>3</sub>F coordination polyhedron corresponds to a distorted trigonal pyramid. The deviation of the silicon atom from the plane of the equatorial C atoms varies within a fairly narrow range of 0.05 Å. The Si-O bond length decreases along with an increase in Si-F one. These bonds vary in the ranges of 2.135(2)-2.384(1) and 1.681(2)-1.646(1) Å, respectively. Observed differences in the lengths of the Si-O bond are mainly related to intermolecular interactions rather than to the nature of R and R' substituents. The presence of strong N-H...F bonds causes a strengthening of the coordination bond Si-O. In contrast, the formation of a N-H...O bond with the endocyclic O atom leads to weakening of the Si-O coordination bond.

In the solid state, molecules **a** and **e** form chains via N-H...F and N-H...O bonds, respectively; in the cases of **c**, **d** and **f** complexes, centrosymmetric dimers are formed via N-H...O bonds with carbonyl (**d** and **f**) and sulfonyl (**c**) groups of neighboring molecules. The

shortest Si-O bond is observed in the structure of **e**, where F atom is involved in N-H...F hydrogen bond, which leads to the weakening of Si-F bond. In turn, the longest Si-O bond is found in the structure of **b** (2.384(1) Å), which is the maximum value for all the earlier described complexes with OSiC<sub>3</sub>F coordination polyhedron; the reason for this being is the intramolecular N-H...O hydrogen bond. Quantum chemical calculations of isolated molecules and their dimers were carried out to evaluate the energy of hydrogen bonds.

The study was supported by RFBR (grant no. 09-03-00669). The authors are thankful to Prof. Yu. I. Baukov's group for obtaining the single crystals of the complexes analyzed.



**Keywords:** N-organosulfonyl-2-aminoacids, hypervalent silicon, crystal packing

## MS61.P11

*Acta Cryst.* (2011) **A67**, C607-C608

**Crystal structure of [Cu<sub>2</sub>(L)<sub>2</sub>(H<sub>2</sub>O)<sub>2</sub>](SO<sub>4</sub>)<sub>2</sub>·2H<sub>2</sub>O (L = pyridoxal thiosemicarbazone)**

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Dark red monocystals of the title complex were prepared by the reaction of methanol solutions of CuSO<sub>4</sub>·5H<sub>2</sub>O and pyridoxal thiosemicarbazone (L) in mole ratio 1:1. Dinuclear complex cation has centrosymmetric structure. The Cu(II) is situated in a square-pyramidal environment ( $\tau = 0.039$ ). The equatorial plane is formed by chelate ONS coordination of L and one water molecule, while the apical position is occupied by an alcoholic oxygen O(2)<sup>i</sup> ( $i = -x, 1-y, -z$ ) of the bridging ligand. The same way of coordination was reported for a similar ligand [1]. The Cu atom is shifted towards the apical oxygen atom by 0.173(12) Å. The Cu-atom ligand bond lengths are in the range 1.917(2)-1.967(2) Å for O(1) N(3) and O(7), while Cu-S(1) and Cu-O(2)<sup>i</sup> are longer (2.2998(9) Å and 2.233(2) Å, respectively). Pyridoxal thiosemicarbazone is coordinated in neutral, zwitterionic form as tetradentate bridging ligand, *i.e.* via phenol oxygen, hydrazine nitrogen, thioamide sulfur and hydroxymethyl oxygen, which, as the bridging ligand, connects the subunits. Such coordination results in formation of two metalocycles, five-membered (thiosemicarbazide) and six-membered (pyridoxilidene). The coordinated ligand slightly deviates from planarity, and dihedral angles between pyridine ring, six- and five-membered metalocycles are 7.54(10)°, 4.80(7)°, respectively. The six-membered metalocycle is in screw-boat conformation and can be described by the following puckering parameters:  $Q = 0.200(2)$  Å,  $\theta = 67.9(9)^\circ$ ,  $\phi = 34.3(9)^\circ$ , while the other rings are planar. Crystal structure of the complex is stabilized by an extended inter- and intramolecular 3D hydrogen-bond network. It can be mentioned that all possible hydrogen donors are involved in hydrogen-bonding, and that sulfate anion acts as multiple hydrogen acceptor.

The data were collected on an Oxford Diffraction Gemini S diffractometer using MoK $\alpha$  radiation ( $\lambda = 0.71069$  Å), and were corrected for Lorentz, polarization and background effects. The structure was solved by direct methods using SIR92 and refined by full matrix least square methods on  $F^2$  using SHELXL-97 to  $R = 0.036$ . Crystallographic data: C<sub>18</sub>H<sub>32</sub>Cu<sub>2</sub>N<sub>8</sub>O<sub>16</sub>S<sub>4</sub>, Mr = 871.90, monoclinic, space group  $P2_1/n$ ,  $a = 7.2529(2)$  Å,  $b = 11.1766(4)$  Å,  $c = 18.8522(6)$  Å,  $\beta = 93.751(3)^\circ$ ,  $V = 1524.94(8)$  Å<sup>3</sup>,  $Z = 2$ ,  $\mu(\text{MoK}\alpha) = 1.756$  mm<sup>-1</sup>, GooF = 1.059.